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Chemical Reactivity

9th CHEMISTRY  
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CHAPTER NO: 08

**M.C.Q’s**

Q1: Multiple Choice Question:

1. Metals can form ions carrying charges:
   (a) Uni-positive  
   (b) Di-positive  
   (c) tri-positive  
   (d) all of them

2. Which one of the following metal burns with a brick red flame?
   (a) sodium  
   (b) iron  
   (c) magnesium  
   (d) calcium

3. Sodium is extremely reactive metal, but it does not react with:
   (a) hydrogen  
   (b) nitrogen  
   (c) sulphur  
   (d) phosphorus

4. Which one of the following is the lightest metal?
   (a) calcium  
   (b) magnesium  
   (c) lithium  
   (d) sodium

5. Pure alkali metals can be cut simply by knife but iron cannot because of alkali metals have:
   (a) Strong metallic bonding  
   (b) Non- metallic bonding  
   (c) Weak metallic bonding  
   (d) Moderate metallic bonding

6. Which of the following is less malleable?
   (a) sodium  
   (b) iron  
   (c) gold  
   (d) silver

7. Metals lose their electrons easily because:
   (a) They are electronegative  
   (b) They are electropositive  
   (c) they have electron affinity  
   (d) good conductors of heat

8. Which one of the following is brittle?
   (a) sodium  
   (b) aluminium  
   (c) selenium  
   (d) magnesium

9. Which one of the following non-metal is lustrous?
   (a) Sulphur  
   (b) Phosphorus  
   (c) Iodine  
   (d) Carbon

10. Non-metals are generally soft, but which one of the following is extremely hard?
    (a) Graphics  
    (b) Phosphorus  
    (c) Iodine  
    (d) Diamond

11. Which one of the following will not react with dilute HCl?
    (a) Sodium  
    (b) Potassium  
    (c) Carbon  
    (d) Calcium

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Short Questions

Q1: Why reactivity of metal increase down the group?
Ans: The reactivity of metal increase downs the group due to increase in the atomic size. The ability of metals to lose electron increases down the group. Hence, their reactivity increases down the group.

Q2: State physical properties of metals?
Ans: Almost all metals are solids:
   (a) They have high melting and boiling point.
   (b) They process metallic luster and can be polished.
   (c) They have high density.
   (d) They are hard.

Q3: Why nitrogen forms compound with alkaline earth metals?
Ans: Nitrogen forms compound with alkaline earth metals directly because they forms stable nitrides when heated with nitrogen.  
Example: \(3\text{Mg} + \text{N}_2 \rightarrow \text{Mg}_3\text{N}_2\)

Q4: Why the second ionization energy of magnesium is higher than the first one?
Ans: The second ionization energy of magnesium is higher than the first one because the removal of electron from Mg ion is difficult as the nuclear charge attracts the remaining electrons strongly. The size also decreases which contributes the increase in 2nd ionization energy.

Q5: How oxygen react with group 2nd A metals?
Ans: Oxygen reacts with 2nd a metal to give their oxides.  
Example: \(2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}\)

Q6: What is relationship between electropositivity and ionization energy?
Ans: The electropositivity of an element depends on the magnitude of its ionization energy. The lower is the ionization energy of an element the higher is the electropositivity of the element.

Q7: Why electropositive decreases from left to right in a period?
Ans: Electropositive character decreases from left to right in a period due to decrease in atomic size and increase in nuclear charge.

Q8: How electro positivity depends upon size and nuclear charge of an atom?
Ans: The electro positivity depends upon the ionization energy which in turn high nuclear charge have high ionization energy. As a result, atoms having high ionization energy are less electropositive.

Q9: Why ionization energies of alkaline earth metals are higher than alkali metals?
Ans: The ionization energies of alkaline earth metals are higher than alkali metals because alkali metals are having one electron in their last shell which is required to remove while alkaline earth metals have two electrons in their last shell which are required to remove.

Q10: Why silver and gold are least reactive?
Ans: Silver and gold are least reactive because they do not lose electrons easily.

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Q11: Can pure gold be used for making ornaments? If not why
Ans: Pure gold cannot be used for making ornaments because gold is too soft to be used as such. Gold is always alloyed with copper, silver or some other metal.

Q12: Why copper is used for making electrical wires?
Ans: Copper is used for making electrical wires because it is good conductor of electricity and can be easily coverted into wires.

Q13: What is the trend of variation in densities of alkali metals?
Ans: Densities of alkali metals decrease down the group.

Q14: Which metals is used for metal work?
Ans: Copper metal is used in metal work because it is easily workable.

Q15: Why magnesium is harder than sodium?
Ans: Magnesium is harder than sodium because magnesium form stronger metallic bond than sodium.

Q16: Why calcium is more electropositive than to magnesium?
Ans: Calcium is more electropositive than magnesium because calcium has greater size than magnesium and calcium has greater ability to lose the electron than magnesium.

Q17: Why ionization energy of Na is less than Mg?
Ans: Na has larger size and lower nuclear charge than Mg. Hence, ionization energy of Na is less than Mg.

Q18: Why ionization energy of Na is more than K?
Ans: Na has smaller size than K. Hence, Na is more ionization energy than K.
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